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Ethylene Mollier Chart

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0.1 1 10 300 400 500 600 700 800 900
1000 1100 1200 1300 Enthalpy [kJ/kg]
Pressure [MPa] T=-100 T=-50 T=0 T=50
T=100 T=150 T=0 T=-50 T=-90 T=-80
T=-70 T=-60 T=-40

Ethylene Pressure-Enthalpy Diagram - NTNU

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4.88 pressure-enthalpy diagram for ethylene 10 100 1000 10000 700 750 800 850 900 950 1000 1050 1100 1150 1200 enthalpy, Btu/lb pressure (psia) 160 F 140 F 120 F 100 F 80 F 60 F 40 F 20 F 0 F 20 F 240 F 200 F 100 F 0 F 100 F 200 F 300 F liquid liquid-vapor two

phase vapor psia = pounds per square inch absolute

4.88 pressure-enthalpy diagram for ethylene

Pressure-Enthalpy Diagram for Ethylene . So how did we get these points? Points 1 and 3: From the given information in the table, we know that 1 is saturated vapor at -100°C, and 3 is saturated liquid at -25°C. Saturated liquids and vapors are on the liquid-vapor dome (shown in purple). To find points 3 and 1, we follow the isotherms for -100 ...

Ethylene Diagram - MIT

The state of air with dry bulb temperature 21 o C and relative humidity 50% is indicated in the Mollier chart below: From the Mollier chart the

enthalpy of the moist air is approximately 40 kJ/kg. The wet bulb temperature is approximately 14.5 °C. The specific humidity is approximately 0.0076 kg/kg. Alternatively - if the dry bulb temperature is 19 °C and the wet bulb temperature is 12 °C - the state of the air can be visualized in the Mollier chart like:

Mollier Diagram - Engineering ToolBox

Ethylene is a gas at standard conditions. However, at low temperature and/or high pressures the gas becomes a liquid or a solid. The ethylene phase diagram shows the phase behavior with changes in temperature and pressure. The curve between the critical point and the triple point shows the ethylene boiling point

with changes in pressure.

Ethylene - Thermophysical Properties

Pressure-Enthalpy Diagram S=Specific Entropy, kJ/kg*K T=Temperature, °C Ethane Produced by I. Aartun, NTNU 2002. Based on the program Allprops, Center for Applied Thermodynamic Studies, University of Idaho.

100 Ethane Pressure-Enthalpy Diagram S=5.0 S=5.5 S ...

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0.1 1 10-300 -200 -100 0 100 200 300
400 500 600 700 Enthalpy [kJ/kg]
Pressure [MPa] Propylene Pressure-
Enthalpy Diagram S=Specific Entropy,
kJ/kg*K

Propylene Pressure-Enthalpy Diagram

An enthalpy-entropy chart, also known as the H-S chart or Mollier diagram, plots the total heat against entropy, describing the enthalpy of a thermodynamic system. A typical chart covers a pressure range of 0.01-1000

bar, and temperatures up to 800 degrees Celsius. It shows enthalpy H in terms of internal energy U , pressure p and volume V using the relationship $H = U + p V$.

Enthalpy-entropy chart - Wikipedia

Ethylene C 2H 4 0.2964 1.5482 1.2518
1.237 Helium He 2.0769 5.1926 3.1156
1.667 Hydrogen H 2 4.1240 14.307
10.183 1.405 Methane CH 4 0.5182
2.2537 1.7354 1.299 Neon Ne 0.4119
1.0299 0.6179 1.667 Nitrogen N 2
0.2968 1.039 0.743 1.400 Octane C 8H
18 0.0729 1.7113 1.6385 1.044 Oxygen
O 2 0.2598 0.918 0.658 1.395 Propane C
3H 8 0.1885 1.6794 1.4909 1.126 Steam
H

PROPERTY TABLES AND CHARTS (SI UNITS)

Figure #21, shows the basic components and piping of an ethylene-propylene cascade system. This system is designed to condense high pressure ethylene gas, and to store it as a liquid at atmospheric pressure. Ethylene is the low stage refrigerant. The low stage compressor is a two - Stage, non- lubricated ring less type reciprocating compressor.

Low Temperature Refrigeration for IQF Freezers Advanced ...

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Ethylene Mollier Chart 0.1 1 10 300 400 500 600 700 800 900 1000 1100 1200 1300 Enthalpy [kJ/kg] Pressure [MPa] T=-100 T=-50 T=0 T=50 T=100 T=150 T=0 T=-50 T=-90 T=-80 T=-70 T=-60 T=-40 Ethylene Pressure-Enthalpy

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Enthalpy Entropy (h-s) or Mollier Diagram . The Mollier diagram, shown in Figure A-1 , is a chart on which enthalpy (h) versus entropy (s) is plotted. It is sometimes known as the h-s diagram and has an entirely different shape from the T-s diagrams.

Enthalpy Entropy (h-s) or Mollier Diagram | Engineers Edge ...

In addition to rendering all H and S values positive throughout the diagram, this realignment resulted in the properties shown in Fig. 6 for saturated vapour at 120K, namely $H = 636.5 \text{ Jg}^{-1}$ $S = 5.304 \text{ Jg}^{-1}\text{K}^{-1}$ Fig. 7 is the Mollier diagram for ethylene obtained by projecting in the pressure direction that portion of Fig. 6 with entropy values between 3.0 and 6.0 $\text{Jg}^{-1}\text{K}^{-1}$ and enthalpy values between 550 and 1300 Jg^{-1} .