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 in the space provided. 1. 88% The actual
 yield of a reaction is 22 g and the
 theoretical yield is 25 g. Calculate the
 percentage yield. 2. 6.0 mol of N₂ are

mixed with 12.0 mol of H₂ according to
 the following equation: N₂(g) + 3H₂(g) →
 2NH₃(g) N

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 yield of a reaction is 22 g and the
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 mixed with 12.0 mol of H₂ according to
 the following equation: N₂(g) + 3H₂(g) ...

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 the following questions in the space
 provided. 1. Given the following
 equation: C₃H₄(g) + x O₂(g) → 3CO₂(g)
 + 2H₂O(g) a. What is the value of the
 coefficient . x. in this equation? b. What
 is the molar mass of C₃H₄? c. How many
 moles are in an 8.0 g sample of C₃H₄?

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 REVIEW SHORT ANSWER Answer the
 following questions in the space
 provided. 1. Given the following
 equation: C₃H₄(g) + xO₂(g) → 3CO
 2(g) + 2H₂O(g) 4 a. What is the value of
 the coefficient x in this equation? 40.07
 g/mol b. What is the molar mass of C₃H
 4? 2 mol O₂:1 mol H₂O

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 22 g and the theoretical yield is 25 g
 Calculate the percentage yield
 2. 60 mol of N₂ are mixed with 120 mol of H₂
 according to the following equation:

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 following questions in the space
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 equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$
 a. What is the value of the coefficient x in this equation? 40.07
 g/mol b.

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 a. What is the value of the coefficient . x. in this equation? b. What
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 moles are in an 8.0 g sample of C₃H₄? 2.
 a. What is meant by . ideal conditions

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Stoichiometry b. Theoretically, how
 many moles of NH₃ will be produced?
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 to the left, Show all your work in the
 space provided. 1. 88% The actual yield
 of a reaction is 22 g and the theoretical
 yield is 25 g. Calculate the percentage
 yield. 2. 6.0 mol of N₂ are mixed with

12.0 mol of H₂ according to the ...

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Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of

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