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Chapter 14 Chemical Equilibrium

Quiz

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At elevated temperatures ammonium carbonate, $\text{NH}_4\text{COONH}_4$, is in equilibrium with NH_3 and CO_2 according to the equation;

$$\text{NH}_4\text{COONH}_4(s) \rightleftharpoons 2\text{NH}_3(g) + \text{CO}_2(g)$$

What is the equilibrium expression for this reaction?

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Example 14.2 Manipulating the Equilibrium Constant to Reflect Changes

in the Chemical Equation Consider the chemical equation and equilibrium constant for the synthesis of ammonia at 25 °C: $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightleftharpoons 2 \text{NH}_3(\text{g})$ $K = 3.7 \times 10^8$ Calculate the equilibrium constant for the following reaction at 25 °C: $\text{NH}_3(\text{g}) \rightleftharpoons \frac{1}{2} \text{N}_2(\text{g}) + \frac{3}{2} \text{H}_2(\text{g})$ $K' = ?$
SOLUTION

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Chem 1120 - Chapter 14: Chemical Equilibrium Practice Quiz 1. 1. The value of the equilibrium constant for the reaction: $2 \text{HBr}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{Br}_2(\text{g})$ is $K_c = 1.26 \times 10^{-12}$ at 500 K. This implies that: a) the product concentrations will be large relative to the reactants at equilibrium. b) the reaction has a large negative ΔG° .

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===== Chapter 14: Chemical

Equilibrium [1] Which is the correct equilibrium constant expression for the following reaction? $\text{Fe}_2\text{O}_3(\text{s}) + 3 \text{H}_2(\text{g}) \rightleftharpoons 2 \text{Fe}(\text{s}) + 3 \text{H}_2\text{O}(\text{g})$ A) $K_c = [\text{Fe}_2\text{O}_3] [\text{H}_2]^3 / [\text{Fe}]^2 [\text{H}_2\text{O}]^3$ D) $K_c = [\text{Fe}]^2 [\text{H}_2\text{O}]^3 / [\text{Fe}]$

CHAPTER-14

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Chapter 14: Chemical Equilibrium Q1. A reaction with an equilibrium constant $K_c = 1.5 \times 10^{21}$ would consist of which of the following at equilibrium: A) approximately equal reactants and products B) some reactants and products with reactants slightly favored

Chapter 14: Chemical Equilibrium

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Chapter 14 Equilibrium Notes page 1 of 6 Chapter 14. CHEMICAL EQUILIBRIUM 14.1 THE CONCEPT OF EQUILIBRIUM AND THE EQUILIBRIUM CONSTANT Many chemical reactions do not go to completion but instead attain a state of chemical equilibrium. Chemical

equilibrium: A state in which the rates of the forward and reverse reactions ...

Chapter 14. CHEMICAL EQUILIBRIUM

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A) It is a process used for shifting equilibrium positions to the right for more economical chemical synthesis of a variety of substances. B) It is a process used for the synthesis of ammonia. C) It is another way of stating Le Châtelier's principle. D) It is an industrial synthesis of sodium chloride that was discovered by Karl Haber.

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14) Consider the following reaction at equilibrium: $2\text{CO}_2(\text{g}) \rightleftharpoons 2\text{CO}(\text{g}) + \text{O}_2(\text{g})$
 $\Delta H^\circ = -514 \text{ kJ}$ Le Chatelier's principle predicts that adding $\text{O}_2(\text{g})$ to the

reaction container will _____. A) decrease the partial pressure of $\text{CO}_2(\text{g})$ at equilibrium B) decrease the value of the equilibrium constant

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