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Chapter 9 Stoichiometry Worksheet

9.1. Objectives.

- Define stoichiometry.
 - Describe the importance of the mole ratio in stoichiometric calculations.
 - Write a mole ratio relating two substances in a chemical equation.
- 2/26/2018
2. Stoich Definitions.
- Composition
- stoichiometry deals with the mass relationships of elements in compounds.
- Reaction
- stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

Chapter 9

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on

CHAPTER 9

Stoichiometry - Quia

Stoichiometry Chapter 9, p. 275 - 294 Intro to Stoichiometry •

Reaction

Stoichiometry: Involves the mass relationships between reactants and products in a chemical reaction • Coefficients in a chemical reaction represent the mole ratios of each substance that react together • Example:



Chapter 9

Stoichiometry Notes

9.3 Objectives Describe a method for

determining which of two reactants is a limiting reactant.

Calculate the amount in moles or mass in

grams of a product, given the amounts in moles or masses in

grams of two reactants, one of which is in excess.;

Distinguish between theoretical yield, actual yield, and percentage yield.; Calculate percentage yield, given the actual yield and

Chapter 9: Stoichiometry - Chem Home

Chapter 9 Worksheet
Stoichiometry 1.

Consider the following equation: $\text{As}_2\text{O}_3 + 2\text{Cl}_2 + 5\text{H}_2\text{O} \rightarrow 2\text{H}_3\text{AsO}_4 + 4\text{HCl}$

HCl a. How many moles of H_2O can be produced from 3.37 mol of H_2O ? b. How many grams of H_2O are needed to react with 0.789 mol of Cl_2 ? c. How many moles of As_2O_3 are needed to produce 62.7g of HCl? d.

Solved: Chapter 9

Worksheet

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Stoichiometry

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CHAPTER 9 REVIEW

Stoichiometry SECTION

3 PROBLEMS Write the
answer on the line to
the left. Show all your

work in the space
provided. 1. 88% The
actual yield of a
reaction is 22 g and
the theoretical yield is
25 g. Calculate the
percentage yield. 2.
6.0 mol of N₂ are
mixed with 12.0 mol of
H₂ according to the
following equation: N₂(g) + 3H₂(g) → 2NH₃(g)

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Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide.
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Chapter 9 Stoichiometry Section 2 Worksheet

Fermentation

Worksheets - there are 8 printable worksheets for this topic.

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Fermentation Worksheets - Teacher Worksheets

Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known

reactants or products.

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