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sentence or statement
is true or false. T F. 1.
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generally absorb more
energy than chemical
changes. T F. 2. In
exothermic reactions,
the reactants have
more kinetic energy
than the products.

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amount of energy
absorbed by 1.00 g of
water, $\text{H}_2\text{O(l)}$, when its
temperature is raised

10.5 °C is 43.9 J. 36.

Given: V HCl(aq)

=100.0 mL; V

NaOH(aq) =100.0 mL;

T initial =23.5 °C; T

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5 Quiz. True/False.

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sentence or statement
is true or false. 1.

Nuclear changes
generally absorb more
energy than chemical
changes. 2. In
exothermic reactions,
the reactants have
more kinetic energy
than the products. 3.

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increases by a factor of
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solution from 20.5 °C
to 26.0 °C is an
exothermic reaction.

(d) The dissolution of a
salt, which changes the
temperature of the
solution from 21 °C to
19 °C is an
endothermic reaction.
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