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 Aluminum reacts with
 oxygen to produce
 aluminum oxide as

follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate: $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$...

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 12: Stoichiometry Mole
 Ratios Questions 1.
 Aluminum reacts with
 oxygen to produce
 aluminum oxide as
 follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ a. If you use
 2.3 moles of Al, how
 many moles of Al_2O_3
 can you make? b. If
 you want 3.9 moles of
 Al_2O_3 , how many
 moles of O_2 are
 needed? 2. Chemistry
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Answer: 4.93×10^{-5} L or 49.3 μ L In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as ...

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16; Chapter 17;

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The instantaneous rate

is the rate of a reaction

at any particular point

in time, a period of

time that is so short

that the ...

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- Chapter 11 -

Stoichiometry Section

11.1 What is

stoichiometry? 1. true

2. true 3. false 4. true

5. true 6. 2, 2, 64.10 7.

3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

methanol and oxygen

gas 11. carbon dioxide

and water 12. 160.10 g

13. 160.10 g 14. They

are equal. 15. A mole ratio is a ratio between the numbers of moles of any two substances in a ...

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