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3+. (i) The energy level diagram showing the electrons in the five 3d orbitals of a chromium atom is represented below. Draw the completed diagram showing the d orbitals in [Cr(H₂O)₆]³⁺ after splitting. [1] E (ii) State and explain what happens to the splitting of the d orbitals if the ligand is changed from H₂O to NH₃.

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