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Acid/Base Dissociation Constant

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 acid. $\text{H}_2\text{SO}_4 \rightleftharpoons \text{HSO}_4^- + \text{H}^+$. $K_1 = 2.4 \times 10^{-6}$
 $K_2 = 1.0 \times 10^{-2}$ 6.62 1.99 . chromic
 acid. $\text{H}_2\text{CrO}_4 \rightleftharpoons \text{HCrO}_4^- + \text{H}^+$. $K_1 = 3.55$ K_2
 $= 3.36 \times 10^{-7}$ 0.55 6.47 . sulfurous acid.

H_2SO_3 HSO_3^- $K_1 = 1.71 \times 10^{-2}$ $K_2 = 5.98 \times 10^{-8}$.
 H_3PO_4 H_2PO_4^- HPO_4^{2-} $K_1 = 7.1 \times 10^{-3}$ $K_2 = 6.2 \times 10^{-8}$

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For many practical purposes it is more convenient to discuss the logarithmic constant, pK_a. pK_a = -log₁₀ K_a. The larger the value of pK_a, the smaller the extent of dissociation at any given pH - that is, the weaker the acid. A weak acid has a pK_a value in the approximate range -2 to 12 in water.

Inorganic acids and bases - pK_a values

Acid Ionization Constants at 25°C; Name of Acid Ionization Equation K_a; Sulfuric acid: $\text{H}_2\text{SO}_4 \rightleftharpoons \text{H}^+ + \text{HSO}_4^-$ $\text{HSO}_4^- \rightleftharpoons \text{H}^+ + \text{SO}_4^{2-}$ very large. 1.3×10^{-2} . Oxalic acid: $\text{H}_2\text{C}_2\text{O}_4 \rightleftharpoons \text{H}^+ + \text{HC}_2\text{O}_4^-$ $\text{HC}_2\text{O}_4^- \rightleftharpoons \text{H}^+ + \text{C}_2\text{O}_4^{2-}$ 6.5×10^{-2} . 6.1×10^{-5} . Phosphoric acid: $\text{H}_3\text{PO}_4 \rightleftharpoons \text{H}^+ + \text{H}_2\text{PO}_4^-$ $\text{H}_2\text{PO}_4^- \rightleftharpoons \text{H}^+ + \text{HPO}_4^{2-}$ $\text{HPO}_4^{2-} \rightleftharpoons \text{H}^+ + \text{PO}_4^{3-}$ 7.5×10^{-3} . 6.2×10^{-8}

Acid and Base Ionization Constants | Chemistry for Non-Majors

Enthalpy and entropy of ionization in organic solvents are compared with aqueous systems. The impact of the solvent on the ionization constants is interpreted based on the free energy of

transfer applied to all particles involved in the ionization reaction of acids and bases, and the concept of the 'medium effect' on these species.

Ionization constants of weak acids and bases in organic ...

An acid dissociation constant, K_a , (also known as acidity constant, or acid-ionization constant) is a quantitative measure of the strength of an acid in solution. It is the equilibrium constant for a chemical reaction $\text{HA} \rightleftharpoons \text{H}^+ + \text{A}^-$ known as dissociation in the context of acid-base reactions. The chemical species HA is an acid that dissociates into A^- , the conjugate base of the ...

Acid dissociation constant - Wikipedia

The first measurements of the ionization constants of water and aqueous acids and bases were made by Noyes (1907), who used the change in conductance associated with ionization to measure equilibrium constants up to about 300 °C at steam saturation.

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Ionization Constants of Inorganic Acids and Bases in Aqueous Solution, Second Edition D. D. Perrin (Compiler), Pergamon Press, Inc.. Elmsford, NY, 1982. xiii + 180 pp. Tables. 19 X 28 cm. \$50.00. The compilation of tables in this volume is apart of the continuing work of the Commission on Equilibrium Data, Analytical

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Because of the very large range of acid strengths (greater than 10^{-40}), a logarithmic scale of acidity (pK_a) is normally employed. Stronger acids have smaller or more negative pK_a values than do weaker acids. A discussion of acid-base terminology is available here. The pK_a values given here are extrapolated for water at 25 °C. Many of the pK_a values given for weak carbon acids are ...

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