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Chemistry

CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a.

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 following questions in
 the space provided. 1.
 Given the following
 equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ ____ a. What
 is the value of the
 coefficient x in this
 equation? ____ b.

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 CHAPTER 9 REVIEW
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 ANSWER Answer the
 following questions in
 the space provided. 1.
 Given the following
 equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) +$

$2H_2O(g)$ 4 a. What is
 the value of the
 coefficient x in this
 equation? 40.07 g/mol
 b.

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Chapter 9 -
 Stoichiometry 9-1
 Introduction to
 Stoichiometry
 Composition
 Stoichiometry - deals
 with mass relationships
 of elements in
 compounds Reaction
 Stoichiometry -
 Involves mass
 relationships between
 reactants and products
 in a chemical reaction
 I. Reaction
 Stoichiometry
 Problems A. Four
 problem Types, One
 Common Solution

**Chapter 9 -
 Stoichiometry**
 CHAPTER 9 REVIEW

Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...

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PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N₂ are mixed with

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Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of

conservation of mass.
All reaction
stoichiometry ... 290
Chapter 9 DO NOT
EDIT--Changes must be
made through "File
info" ... The number of
significant figures in
the answer

**CorrectionKey=NL-A
DO NOT EDIT--
Changes must be
made ...**

Stoichiometry 4
Chapter 9 Assignment
& Problem Set 5.
Isopropyl alcohol ($\text{C}_3\text{H}_7\text{OH}$) burns in air
according to this
equation: $2\text{C}_3\text{H}_7\text{OH(l)} + 9\text{O}_2\text{(g)} \rightarrow 6\text{CO}_2\text{(g)} + 8\text{H}_2\text{O(g)}$ a.
Calculate the moles of
oxygen needed to
react with 3.40 mol C_3

H_7OH . b. Find the
moles of each product
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 $\text{C}_3\text{H}_7\text{OH}$ reacts with
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