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Section 10 1 The Mole

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SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

Section 10.1 – The Mole: A Measurement of Matter You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is 6.02 x 1023representative particles of that substance.

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Section 10.1 • Measuring Matter 321 Figure 10.2 The amount of each substance shown is 6.02 × 1 0 23 or 1 mol of repre-sentative particles. The representative particle for each substance is shown in a box. Refer to Table R-1 on page 968 for a key to atom color conventions. The mole The mole, abbreviated mol, is the SI base unit used to

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Section 10.1 The Mole: A Measurement of Matter 289 Practice Problems Practice Problems withChemASAP Math Handbook SAMPLE PROBLEM 10.1 Finding Mass from a Count What is the mass of 90 average-sized apples if 1 dozen of the apples has a mass of 2.0 kg? Analyze List the knowns and

the unknown. Knowns • number of apples 90 apples • 12 apples 1 dozen apples

10.1 The Mole: A Measurement of Matter 10

One mole(mol) of a substance is 6.02 × 1023. representative particles of that substance and is the SI unit for measuring the amount of a substance. The number of representative particles in a mole, 6.02 × 1023, is called Avogadro’s number. 10.1. 2/21/15 6.

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Chemists use the term mole to represent a large number of atoms or molecules. Just as a dozen implies 12 things, a mole (mol) represents 6.022 × 10 23 things. The number 6.022 × 10 23, called Avogadro’s number after the 19th-century chemist Amedeo Avogadro, is the number we use in chemistry to represent macroscopic amounts of atoms and ...

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particles: 1.25 1025 atoms of zinc (Zn), 3.56 mol of iron (Fe), and 6.78 1022 mole-cules of glucose (C 6 H 12 O 6). 3.56 mol Fe 6.02 1023 atoms Fe __ 1 mol Fe 24 2.14 10 atoms Fe From smallest to largest: 6.78 1022 molecules glucose, 2.14 1024 atoms Fe, 1.25 1025 atoms Zn. Section 10.2 Mass and the Mole page 325–332 Problem-Solving Lab page 326

The MoleThe Mole - Weebly

A mole is \{(6.022 \times 10^ {23})\} things. Notice that we are applying the mole unit to different types of chemical entities.The word mole represents a number of things—6.022 × 10 23 of them—but does not by itself specify what “they” are. The chemical entities can be atoms, molecules, formula units and ions.This specific information needs to be specified accurately.