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Mnemonic Device For Thermodynamic Potentials and Maxwell's Relations

Thermo: Lesson 1 - Intro to Thermodynamics

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 some of the most
 common conversion
 factors you need in
 thermodynamics.
 Acceleration: 1 m/s² =
 100 cm/s². Area: 1 m² =
 10⁴ cm² = 10⁶ mm². Density: 1 g/cm³ =
 1 kg/L = 1,000 kg/m³. Energy, heat,
 work, internal energy,
 enthalpy: 1 kJ = 1,000 J
 = 1,000 N·m = 1

kPa·m 3. 1 kJ/kg =
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Basic Thermodynamic
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Control Mass (no mass
flow across system
boundaries)

Conservation of mass:

$$\dot{m} = \dot{m}_{in} - \dot{m}_{out}$$

Conservation of energy

(1st Law): $\dot{Q} - \dot{W} = \Delta \dot{E} =$

$$\Delta \dot{E} + \Delta \dot{E}_{kin} + \Delta \dot{E}_{pot} = \dot{Q} - \dot{W}$$

$$\dot{Q} - \dot{W} = \dot{Q}_{in} - \dot{Q}_{out} + \dot{W}_{in} - \dot{W}_{out}$$

$$\dot{Q}(\dot{Q}_{in} - \dot{Q}_{out})$$

Basic

Thermodynamic Formulas (Exam Equation Sheet)

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Thermodynamics key

facts (7/9) • Ideal gas law • 1. st. form : $pV = nRT$ • p = Pressure, V = Volume, n = number of molecules, R = Boltzmann's constant, T = temperature [in K] • 2. nd. form : $p = \frac{1}{3} n \overline{mv^2}$

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Heat & thermodynamics

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Thermodynamics key facts (3/9) • Heat energy needed to change phase • The latent heat L determines the energy Q needed to change the phase of a mass m . • Units of L will be J/kg . – 1. - can be fusion or vaporization • This energy is either absorbed (solid → liquid → gas) or released (gas → liquid → solid) • A phase change takes place at constant temperature.

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Engineering Thermodynamics

Cheat Sheet

Thermodynamics ' ΔT A
 $v = A^2 v P$ = rate of
heat transfer $Q =$
thermal energy $A =$
Area of thermal
conductivity $U =$
coefficient of heat
conductivity (U-factor)
 $T =$ change in
temperature $R =$
resistance to heat flow
(R-value) $v =$ velocity
net = net power
radiated = $5.6696 \times$
 $10^{-8} \frac{1}{\text{ft}^2 \cdot \text{h} \cdot ^\circ\text{F}}$, $T_2 =$
temperature at time 1,
time 2 $v =$ flow velocity

Engineering Formula Sheet

IV. The First Law of
Thermodynamics as a
Rate Equation A. Most
general form $dE_{c.v.}/dt$
 $= \dot{Q}_{c.v.} - \dot{W}_{c.v.} + \dot{m}_e e_{in} - \dot{m}_o e_{out}$
rate of change of
energy in c.v. = rate of
heat added to c.v. rate
of work done + rate of
energy flow in to c.v.

rate of energy flow out of c.v. B. For a steady flow process $d/dt = 0$ and $\dot{m}_{in} = \dot{m}_{out} = \dot{m}$ $\dot{Q}$ c.v. $\dot{W}$

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Thermodynamics

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Figure 2: polytropic process diagrams ...

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Thermodynamics: the

study of energy, energy transformations and its relation to matter. The analysis of thermal systems is achieved through the application of the governing conservation equations, namely Conservation of Mass, Conservation of Energy (1st law of thermodynamics), the 2nd law of thermodynamics and the property relations.

Basic Concepts of Thermodynamics

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Thermodynamics Equations Cheat Sheet

engineering work, pressures are often measured with respect to atmospheric pressure rather than with respect to absolute vacuum. $P_{abs} = P_{atm} + P_{gauge}$
In SI units the derived unit for pressure is the Pascal (Pa), where $1 \text{ Pa} = 1 \text{ N/m}^2$. This is very small for engineering purposes, so usually pressures are given in terms of kilopascals ($1 \text{ kPa} = 1000 \text{ Pa}$),