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where k k k is the rate constant and n n n is the order of the reaction.; Half life of a reaction, $t_{1/2}$ $t_{1/2}$, is the time it takes for one-half of a quantity of a reactant to react. According to the question the time taken for the concentration of the reactant to decrease from $[A]_0$ to $[A]_0/2$ is the same as the time taken to decrease the concentration from $[A]_0/2$ to $[A]_0/4$.

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Chapter 17: ReactionRates

Ans. $\int y = 1 \, 3 \, m \int 2 \, m = r \, A \int = 1 \, 3 \, r \, A \int 3 = L \int 0 \times 2 \, (r \, A \, dx)$

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Chapter 16: Reaction Rates

CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS 17.1 If the rate of the forward reaction exceeds the rate of reverse reaction, products are formed faster than they are consumed. The change in reaction conditions results in more products and less reactants.