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Answer: 4.93 x 10 –5 L or 49.3 µL In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: [Au(CN) 2] – , LaCl 3 , ethanol, and para -nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as ...

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Since 5 pickles combine with 2 ham slices in each sandwich, 20 pickles are needed to fully combine with 8 ham slices. Step 3: Think about your result. The 8 ham slices will make 4 ham sandwiches. With 5 pickles per sandwich, the 20 pickles are used in the 4 sandwiches.

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Chapter 12 Stoichiometry Section 12.1 The Arithmetic of Equations Using Balanced Chemical Equations Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. Stoichiometry = calculation of quantities in chemical reactions is a subject of chemistry.

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Stoichiometry is... Greek for “measuring elements” Defined as: calculations of the quantities in chemical reactions, based on a balanced equation. There are 4 ways to interpret a balanced chemical equation 2

Essential question: What are the 4 ways to interpret a ...

Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall that the law states that matter is neither created nor destroyed in a chemical reaction.

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