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The Mole Webquest Answer Key

5) How many grams of water are in one mole of water? 1 mole water = 18 g water 6) How many molecules of water are in one mole of water? 6.02×10^{23} Use the following website to answer the next set of questions: 7) Define molar mass.

Mole Webquest.docx answer key.docx - Name Date The Mole ...

a. What is the atomic weight of one mole of CH₄? b. How many moles in 64 grams of CH₄? 10) What is the formula to determine the mass in grams of 3.6 mols of H₂SO₄? 11) YOU TRY (not on website) a. What is the mass in grams of 4.2 mols of FeO₂? Use

the following website to answer the next set of questions:

The Mole Webquest - Eldon Middle School

Begin at one of the two starting locations (Start 1 and 2) and work through the questions to reach the finish space. In each space there is a question that continues from the statement in the start box. If the content in the space equals one mole, follow the "T" for true arrow. If it does not equal one mole, follow the "F" for false arrow.

The Mole Webquest - Chemistry by Mrs. Hinkson

5. What is the weight of 4.30 mole of sodium? 6. What is the weight of 1.75 mole of Ca? 7. How many moles are there in 85.3 g of P? 8. Change 0.566 g of silver to moles. Gram-mole Practice Problems: 1. Change 3.4 moles of HCl to grams. 2. Change 8.5 g of SiH₄ to moles. 3. Change 5.20 moles of C₃H₆O₂ to grams. 4. Change 13.2 g of Fe(NO₃)₃ ...

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Mole Webquest Answer Key

Since a molecular weight in grams (mole) of any substance contains the same number of molecules, then according to Avogadro's Principle, the molar volumes of all gases should be the same. The number of molecules in one mole is now called

The Mole Webquest - Huber Heights City Schools

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The number of representative particles in one mole of a substance 6.02×10^{23} . Converting from atoms to moles. Divide the number of particles by Avogadro's number to determine the number of moles. ... ANSWER: Example converting moles to

formula units. QUESTION: Example converting grams to molecules. ANSWER: Example converting grams to ...

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5.1 Balance chemical equations by applying the laws of conservation of mass and constant composition (definite proportions). 5.3 Use the mole concept to determine number of particles and molar mass for elements and compounds. 5.4 Determine percent compositions, empirical formulas, and molecular formulas.

Stoichiometry | Create WebQuest

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